

Low-cost battery from waste graphite

October 11 2017, by Rainer Klose

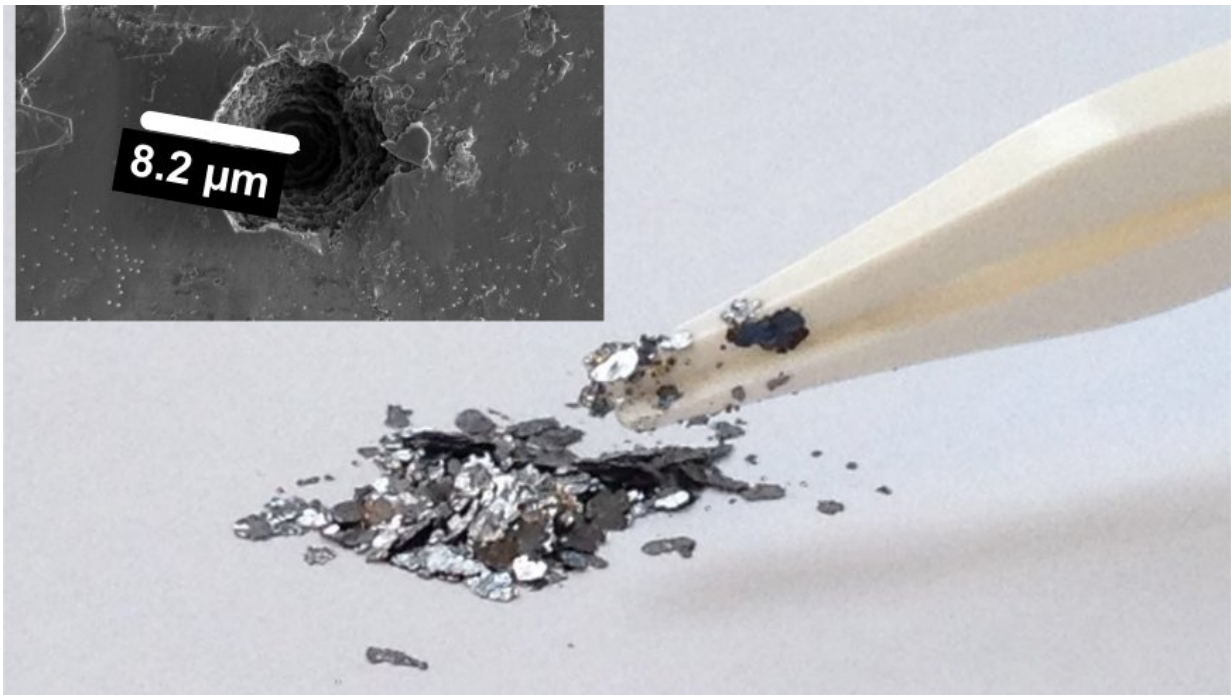


Kostiantyn Kravchyk explains, why kish graphite works so good: The edges of the graphite look like the edges of a bunch of paper. The ions can move in there easily. If the graphite layers are crunched, that won't work so good. Credit: Empa

Lithium ion batteries are flammable and the price of the raw material is rising. Are there alternatives? Yes: Empa and ETH Zürich researchers have discovered promising approaches as to how we might produce batteries out of waste graphite and scrap metal.

Kostiantyn Kravchyk works in the group of Maksym Kovalenko. This research group is based at both ETH Zurich and in Empa's Laboratory for Thin Films and Photovoltaics. The two researchers' ambitious goal at the Empa branch is to make a [battery](#) out of the most common elements in the Earth's crust – such as magnesium or aluminum. These metals offer a high degree of safety, even if the anode is made of pure metal. This also offers the opportunity to assemble the batteries in a very simple and inexpensive way and to rapidly upscale the production.

In order to make such batteries run, the liquid electrolyte needs to consist of special ions that do not crystallize at room temperature – i.e. form a kind of melt. The metal ions move back and forth between the cathode and the anode in this "cold melt", encased in a thick mantle of chloride ions. Alternatively, large but lightweight organic anions, which are metal-free, could be used. This does come with a problem, though: where are these "thick" ions supposed to go when the battery is charged? What could be a suited [cathode material](#)? By way of comparison: in lithium ion batteries, the cathode is made of a metal oxide, which can easily absorb the small lithium cations during charging. This does not work for such large ions, however. In addition, these large anions have an opposite charge to the lithium cations.



Kish graphite is a waste product from steel production. It could be used to make a cheap rechargeable battery out of abundant materials. Credit: Empa / ETH Zürich

To solve the problem, Kovalenko's team had a trick up their sleeves: the researchers turned the principle of the [lithium ion battery](#) upside down. In conventional Li-ion batteries, the anode (the negative pole) is made of [graphite](#), the layers of which (in a charged state) contain the lithium ions. In Kovalenko's battery, on contrary, the graphite is used as a cathode (the positive pole). The thick anions are deposited in-between the graphene layers. In Kovalenko's battery, the anode is made of metal.

Kravchyk made a remarkable discovery while searching for the "right" graphite: He found that waste graphite produced in steel production, referred to as "kish graphite", makes for a great cathode material. Natural graphite also works equally well – if it is supplied in coarse

flakes and not ground too finely or into folded, non-flake shapes. The reason: the graphite layers are open at the flakes' edges and the thick anions are thus able to slip into the structure more easily. The fine-ground graphite normally used in [lithium ion](#) batteries, however, is ill-suited for Kovalenko's battery. By grinding the graphite particles, the layers become creased like crumpled-up paper. Only small [lithium](#) ions are able to penetrate this crumpled graphite, not the new battery's thick anions.

The graphite cathode battery constructed from steel production "kish graphite" or raw, natural graphite flakes has the potential to become highly cost-effective. And if the first experiments are anything to go by, it is also long-lasting. For several months, a lab system survived thousands of charging and discharging cycles. "The aluminum chloride – graphite [cathode](#) battery could last decades in everyday household use," explains Kravchyk and adds "similar demonstrations, but further increased battery voltages, without compromising capacities, and of even lighter elements are on the way and will offer further increase in energy densities from current 60 Wh kg⁻¹ to above 150 Wh kg⁻¹"

More information: Shutao Wang et al. Kish Graphite Flakes as a Cathode Material for an Aluminum Chloride–Graphite Battery, *ACS Applied Materials & Interfaces* (2017). [DOI: 10.1021/acsami.7b07499](https://doi.org/10.1021/acsami.7b07499)

Provided by Swiss Federal Laboratories for Materials Science and Technology

Citation: Low-cost battery from waste graphite (2017, October 11) retrieved 4 October 2025 from <https://phys.org/news/2017-10-low-cost-battery-graphite.html>

This document is subject to copyright. Apart from any fair dealing for the purpose of private

study or research, no part may be reproduced without the written permission. The content is provided for information purposes only.